Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**ATOMIC RADIUS & ION RADIUS TRENDS**

1. In the following, rank the atoms from smallest to largest atomic radius:
   1. O, F, N
   2. K, Rb, Na
   3. K, Ca, Mg
   4. Br, As, S
   5. Si, C, N
   6. Na+, Na
   7. Cl, Cl-1
   8. Be+2, Na+1, Ne
   9. N-3, O-2, F-1, Ne
2. Draw a diagram to help you remember the atomic radii trends and ionic radii trends.
3. Explain why as you go across a period, the atomic radius becomes smaller.
4. Explain why anions are larger and cations are smaller than their parent atom.