Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**ATOMIC RADIUS & ION RADIUS TRENDS**

1. In the following, rank the atoms from smallest to largest atomic radius:
	1. O, F, N
	2. K, Rb, Na
	3. K, Ca, Mg
	4. Br, As, S
	5. Si, C, N
	6. Na+, Na
	7. Cl, Cl-1
	8. Be+2, Na+1, Ne
	9. N-3, O-2, F-1, Ne
2. Draw a diagram to help you remember the atomic radii trends and ionic radii trends.
3. Explain why as you go across a period, the atomic radius becomes smaller.
4. Explain why anions are larger and cations are smaller than their parent atom.