**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_**

**Chapter 2 Review**

**Matter**

1. **What is Matter?**
2. **What is the Kinetic Molecular Theory and what does it say?**
3. **What is the difference between the Law of Definite Composition and the Law of Multiple Proportions?**
4. **What is the difference between pure substances and mixtures?**
5. **What is the difference between an element and a compound?**
6. **What is the difference between a homogeneous mixture and heterogeneous mixture?**
7. **What is the difference between miscible and immiscible?**
8. **What is density?**
9. **I have a chemical compound with the formula KNO3.How many of each element do I have?**
	1. **K-**
	2. **N-**
	3. **O-**
10. **Give an example of an element.**
11. **Give an example of a compound.**
12. **Give an example of a heterogeneous mixture.**
13. **Give an example of a homogeneous mixture.**
14. **How do we know if a chemical reaction has taken place?**
15. **How do we reverse chemical changes?**
16. **What gives elements its properties?**
17. **I have the element Aluminum. Does the compound Aluminum Chloride have the same properties?**
18. **What is the difference between a physical property and a physical change?**
19. **An object of has a mass of 48 g and a volume of 6 mL. What is its density? (D=M÷V)**
20. **An object has a volume of 2.5 L and a density of 8 kg/L. What is its mass? (D=M÷V)**
21. **Liquid A has a density of 4.6 g/mL and Liquid B has a density of 3.1 g/mL which one would be the layer on top if both are placed in a container?**