**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Ionic Practice #3 – Transition Metals**

**Fill in the chart below.**

|  |  |  |  |
| --- | --- | --- | --- |
| Compound Name | **WRITE CHARGES HERE:** | **CAN THE NUMBERS BE REDUCED?** | **CROSS CHARGES TO GET FORMULA:** |
| 1. Copper (I) Chlorine |  |  |  |
| 1. Iron (III) Sulfide |  |  |  |
| 1. Zinc (II) Iodide |  |  |  |
| 1. Tin (IV) Selenide |  |  |  |
| 1. Palladium (IV) Bromide |  |  |  |
| 1. Nickel (II) Carbonate |  |  |  |
| 1. Chromium (III) Sulfate |  |  |  |
| 1. Silver (I) Nitrate |  |  |  |
| 1. Titanium (II) Phosphate |  |  |  |
| 1. Scandium (III) Hydroxide |  |  |  |

**Give the name for the following ionic compounds:**

1. CuSO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. Pb(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Co3(PO4)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. FeBr3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. Sn3(PO4)4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. Mn(NO3)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
7. FePO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
8. CoCO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. CuCl2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
10. Zn(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_