**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Ionic Practice #4 – with Mixed Ionics**

**Fill in the chart below.**

|  |  |  |  |
| --- | --- | --- | --- |
| Compound Name | **WRITE SYMBOL & CHARGES HERE:** | **CAN THE NUMBERS BE REDUCED?** | **CROSS CHARGES TO GET FORMULA:** |
| 1. Potassium Iodide |  |  |  |
| 1. Copper (II) Hydroxide |  |  |  |
| 1. Magnesium Oxide |  |  |  |
| 1. Iron (III) Sulfate |  |  |  |
| 1. Strontium Fluoride |  |  |  |
| 1. Gallium Acetate |  |  |  |
| 1. Tin (IV) Selenide |  |  |  |
| 1. Ammonium Sulfide |  |  |  |
| 1. Aluminum Phosphate |  |  |  |
| 1. Silver (I) Bromide |  |  |  |
| 1. Nickel (III) Hydroxide |  |  |  |
| 1. Sodium Nitrate |  |  |  |
| 1. Palladium (IV) Bromide |  |  |  |
| 1. Beryllium Oxide |  |  |  |
| 1. Aluminum Nitride |  |  |  |

**Give the name for the following ionic compounds:**

1. Na2O\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. LiC2H­3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Ca3(PO4)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. MgCO3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. Al3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. Mg3N2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
7. NH4NO3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
8. NaF \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. Be(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
10. SrS \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
11. NH4I \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
12. Mg(NO3)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
13. KClO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
14. Na2SO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
15. Al(OH)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_