**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Ionic Practice #4 – with Mixed Ionics**

**Fill in the chart below.**

|  |  |  |  |
| --- | --- | --- | --- |
| Compound Name | **WRITE SYMBOL & CHARGES HERE:** | **CAN THE NUMBERS BE REDUCED?** | **CROSS CHARGES TO GET FORMULA:**  |
| 1. Potassium Iodide
 |  |  |  |
| 1. Copper (II) Hydroxide
 |  |  |  |
| 1. Magnesium Oxide
 |  |  |  |
| 1. Iron (III) Sulfate
 |  |  |  |
| 1. Strontium Fluoride
 |  |  |  |
| 1. Gallium Acetate
 |  |  |  |
| 1. Tin (IV) Selenide
 |  |  |  |
| 1. Ammonium Sulfide
 |  |  |  |
| 1. Aluminum Phosphate
 |  |  |  |
| 1. Silver (I) Bromide
 |  |  |  |
| 1. Nickel (III) Hydroxide
 |  |  |  |
| 1. Sodium Nitrate
 |  |  |  |
| 1. Palladium (IV) Bromide
 |  |  |  |
| 1. Beryllium Oxide
 |  |  |  |
| 1. Aluminum Nitride
 |  |  |  |

**Give the name for the following ionic compounds:**

1. Na2O\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. LiC2H­3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Ca3(PO4)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. MgCO3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. Al3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. Mg3N2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
7. NH4NO3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
8. NaF \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. Be(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
10. SrS \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
11. NH4I \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
12. Mg(NO3)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
13. KClO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
14. Na2SO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
15. Al(OH)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_