**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Properties of Ionic & Covalent Bonds**

1. Which type of bonds transfers electrons?
2. Which type of bond shares electrons?
3. What is the difference between a cation and anion?
4. How do neutral atoms become a cation? Use Sodium as an example.
5. How do neutral atoms become an anion? Use chlorine as an example.
6. Show with Lewis Dot Structures how sodium and chlorine form a bond.
7. Show with Lewis Dot Structures how 1 Carbon and 4 bromine form a bond.
8. Explain why NaF is an ionic compound while NF3 is a covalent compound.
9. What charge will the following elements have when they follow the octet rule to become like the nearest noble gas?
10. Bromine \_\_\_\_\_\_
11. Nitrogen \_\_\_\_\_\_
12. Magnesium \_\_\_\_\_\_\_
13. Sulfur \_\_\_\_\_\_
14. Sodium \_\_\_\_\_\_
15. Aluminum \_\_\_\_\_\_
16. Silicon \_\_\_\_\_\_