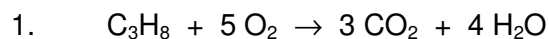


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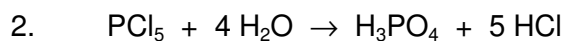
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Chemistry
Stoichiometry WS 1

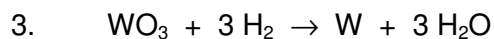
I. *Perform the following stoichiometric calculations:*



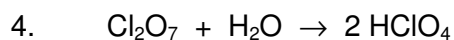
a. 20.2 g of C_3H_8 are burned in the air. How many grams of CO_2 are formed?



a. A chemical reaction requires at least 215 g of HCl. How much PCl_5 must be reacted with water to make this much HCl?



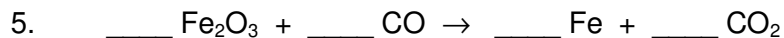
a. How much tungsten is formed from the reaction of 77.28 g of tungsten (VI) oxide?



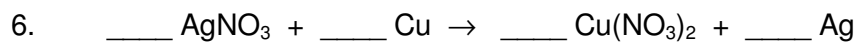
a. In order to react a particular amount of Cl_2O_7 , 95.2 g of H_2O is required. How many grams of Cl_2O_7 can be reacted?

b. In the above problem, how many grams of HClO_4 will be formed by this reaction?

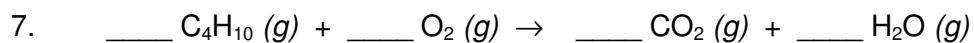
II. Complete the following stoichiometric calculations, balancing equations where necessary:



a. Using the above equation, if 200 g of Fe_2O_3 reacts with an excess of carbon monoxide, how many grams of carbon dioxide are formed?



a. When 11.7 g of copper (II) nitrate are formed in this reaction, how many grams of copper were reacted?



a. 244 g of butane (C_4H_{10}) are burned in excess O_2 . How many grams of CO_2 will be formed?

b. In the above reaction, how many liters of water vapor will be formed at STP?