**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Net Ionic Reaction Practice #1**

1. Identify the reaction type, (b) write the molecular equation with symbols, (c) write the net ionic equation for each of the following:
2. Dilute sulfuric acid is added to a solution of barium acetate.
	1.
	2.
	3.
3. Solid zinc strips are added to a solution of copper (II) sulfate
	1.
	2.
	3.
4. Solid ammonium carbonate is heated
	1.
	2.
	3.
5. Ethanol (C2H5OH) is burned completely in air
	1.
	2.
	3.
6. Solid calcium oxide is heated in the presence of sulfur trioxide gas
	1.
	2.
	3.
7. Lithium metal is burned in air
	1.
	2.
	3.
8. A solution of hydrogen peroxide is heated
	1.
	2.
	3.
9. Sodium chunks are thrown into a pond
	1.
	2.
	3.
10. Equal volumes of 0.1 M sulfuric acid and 0.1 M potassium hydroxide are mixed
	1.
	2.
	3.
11. Solid Iron (II) sulfide is added to aqueous hydrochloric acid
	1.
	2.
	3.