**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Net Ionic Reaction Practice #2**

(a) write the molecular equation with symbols, (c) write the net ionic equation for each of the following:

1. Silver nitrate + potassium chromate
	1.
	2.
2. Ammonium chloride + cobalt (II) sulfate
	1.
	2.
3. Lithium hydroxide + sodium chromate
	1.
4. Zinc acetate + cesium hydroxide
	1.
	2.
5. Ammonium sulfide + Lead (II) nitrate
	1.
	2.
6. Iron (III) sulfate + barium iodide
	1.
	2.
7. Chromium (III) bromide + sodium nitrate
	1.
	2.
8. Rubidium phosphate + titanium (IV) nitrate
	1.
	2.
9. Ammonium carbonate + nickel (II) chloride
	1.
	2.
10. Tin (IV) nitrate + potassium sulfite
	1.
	2.