**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Net Ionic Reaction Practice #3**

(a) write the molecular equation with symbols, (c) write the net ionic equation for each of the following:

All reactants are aqueous unless otherwise stated

1. Ammonium Sulfate + hydrochloric acid
	1.
	2.
2. Ammonium Sulfide + Silver Nitrate
	1.
	2.
3. Cobalt (II) Chloride is combined with silver nitrate
	1.
4. Solid calcium carbonate is reacted with sulfuric acid
	1.
	2.
5. Potassium sulfite is reacted with hydrobromic acid
	1.
	2.
6. Potassium sulfide is reacted with nitric acid
	1.
	2.
7. Ammonium iodide + magnesium sulfate
	1.
	2.
8. Solid titanium (IV) carbonate + hydrochloric acid
	1.
	2.
9. Solid calcium sulfite + acetic acid
	1.
	2.
10. Strontrium hydroxide + ammonium sulfide
	1.
	2.