**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Net Ionic Reaction Practice #4**

(a) write the molecular equation with symbols, (c) write the net ionic equation for each of the following:

1. A sample of calcium carbonate is heated
2. Sulfur dioxide gas is bubbled through water
3. Solid potassium oxide is added to a container of carbon dioxide gas
4. Liquid hydrogen peroxide is warmed
5. Solid lithium oxide is added to water
6. Molten aluminum chloride is electrolyzed
7. A pea sized piece of sodium is added to a container of iodine vapor
8. A sample of carbonic acid is heated
9. A sample of potassium chlorate is heated
10. Solid Magnesium Oxide is added to sulfur trioxide gas