**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Net Ionic Reaction Practice #4**

(a) write the molecular equation with symbols, (c) write the net ionic equation for each of the following:

1. A sample of calcium carbonate is heated
	1.
	2.
2. Sulfur dioxide gas is bubbled through water
	1.
	2.
3. Solid potassium oxide is added to a container of carbon dioxide gas
	1.
4. Liquid hydrogen peroxide is warmed
	1.
	2.
5. Solid lithium oxide is added to water
	1.
	2.
6. Molten aluminum chloride is electrolyzed
	1.
	2.
7. A pea sized piece of sodium is added to a container of iodine vapor
	1.
	2.
8. A sample of carbonic acid is heated
	1.
	2.
9. A sample of potassium chlorate is heated
	1.
	2.
10. Solid Magnesium Oxide is added to sulfur trioxide gas
	1.
	2.