**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Net Ionic Reaction Practice #5**

(a) write the molecular equation with symbols, (c) write the net ionic equation for each of the following:

If there is no reaction, write no rxn.

1. A piece of copper is dropped into a container of water
	1.
	2.
2. Liquid bromine is added to a container of sodium iodide crystals
	1.
	2.
3. An aluminum strip is immersed in a solution of silver nitrate
	1.
4. Zinc pellets are added to a sulfuric acid solution
	1.
	2.
5. Fluorine gas is bubbled into a solution of ammonium chloride
	1.
	2.
6. Magnesium turnings are added to a solution of lead (II) acetate
	1.
	2.
7. Iodine crystals are added to a solution of sodium chloride
	1.
	2.
8. Calcium metal is added to a solution of nitric acid
	1.
	2.
9. A pea-sized piece of lithium is added to water
	1.
	2.
10. A solution of iron (III) chloride is poured over a piece of platinum wire
	1.
	2.