**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Net Ionic Reaction Practice #6**

(a) write the molecular equation with symbols, (c) write the net ionic equation for each of the following:

If there is no reaction, write no rxn.

1. Calcium metal is added to a dilute solution of hydrochloric acid
	1.
	2.
2. Solid sodium oxide is added to water
	1.
	2.
3. Dilute nitric acid is added to a potassium sulfite solution
	1.
4. Calcium chlorate crystals are heated
	1.
	2.
5. Silver nitrate is added to potassium chloride
	1.
	2.
6. magnesium nitrate is added to sodium carbonate
	1.
	2.
7. Rubidium fluoride is added to copper (II) sulfate
	1.
	2.
8. calcium hydroxide is added to iron (III) chloride
	1.
	2.
9. Barium nitrate is added to ammonium phosphate
	1.
	2.
10. Solid potassium chlorate decomposes upon heating
	1.
	2.