**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_**

**Balancing Chemical Reactions Practice #3**

***Directions****: Balance the following reactions. Remember, you have to have the same number of each element on each side of the equation.* ***THEN****, Identify the type of reaction for each problem as Synthesis (S), Decomposition (D), Single Replacement (SR), Double Replacement (DR), or Combustion (C)*

1. \_\_\_\_\_Na(NO3) + \_\_\_\_\_ PbO → \_\_\_\_\_Pb(NO3)2 + \_\_\_\_\_Na2O 1. Type:
2. \_\_\_\_\_Fe + \_\_\_\_\_ Ag(NO3) → \_\_\_\_\_Fe(NO3) + \_\_\_\_\_Ag 2. Type:
3. \_\_\_\_\_Zn(SO4) + \_\_\_\_\_ Li2(CO3) → \_\_\_\_\_Zn(CO3) + \_\_\_\_\_Li2(SO4)*3.* Type:
4. \_\_\_\_\_AgI + \_\_\_\_\_ Fe2(CO3)3 → \_\_\_\_\_FeI3 + \_\_\_\_\_Ag2(CO3) 4. Type:
5. \_\_\_\_\_V2O5 + \_\_\_\_\_ CaS → \_\_\_\_\_CaO + \_\_\_\_\_V2S5 5. Type:
6. \_\_\_\_\_C2H6 + \_\_\_\_\_ O2 → \_\_\_\_\_CO­2 + \_\_\_\_\_H2O 6. Type:
7. \_\_\_\_\_S8 + \_\_\_\_\_ O2 → \_\_\_\_\_SO2 + 7. Type:
8. \_\_\_\_\_AgBr + \_\_\_\_\_ Ga(PO4) → \_\_\_\_\_Ag3(PO4) + \_\_\_\_\_GaBr3 8. Type:
9. \_\_\_\_\_Mn(NO2)2 + \_\_\_\_\_ BeCl2 → \_\_\_\_\_Be(NO2)2 + \_\_\_\_\_MnCl2 9. Type:
10. \_\_\_\_\_KClO3 → \_\_\_\_\_KCl + \_\_\_\_\_O2 10. Type:
11. \_\_\_\_\_Al + \_\_\_\_\_ O2 → \_\_\_\_\_Al2O3 11. Type:
12. \_\_\_\_\_C4H10 + \_\_\_\_\_ O2 → \_\_\_\_\_CO2 + \_\_\_\_\_H2O 12. Type:
13. \_\_\_\_\_H2O → \_\_\_\_\_O2 + \_\_\_\_\_H2 13. Type:
14. \_\_\_\_\_Ca3N2 + +\_\_\_\_\_K → \_\_\_\_\_K3N + \_\_\_\_\_Ca14. Type:
15. \_\_\_\_\_Al + \_\_\_\_\_ HCl → \_\_\_\_\_AlCl3 + \_\_\_\_\_H2 15. Type: