**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_**

**OBSERVING EVIDENCE OF CHEMICAL INTERACTION**

This investigation provides an opportunity to make some careful observations on chemical and physical change. Be alert and pay attention to all the details of the interaction(s) (before, during, and after). A data tables will allow you make both qualitative and quantitative observations.

During this investigation, you will be working with 3 substances which will interact with each other to produce two changes – one chemical, the other physical.

**CAUTION:** Wear your safety glasses at all times. Hair must be in a ponytail. No loose clothing.

**Materials**:

safety glasses aluminum foil mass scale

distilled water copper II chloride glass stir stick

100 mL beaker thermometer

**Procedure:**

1. Obtain a clean 100 mL beaker. Place it on a mass scale. Tare the scale. Carefully add 2.0 g of copper II chloride to the beaker. **Record** the mass of the copper II chloride on the data table 1.
2. **Look** at the colour of copper II chloride (CuCl2) crystals.  **Record** the colour of the crystals on the data table 2.
3. Obtain a 50 mL graduated cylinder. Place 25 mL of distilled water into the graduated cylinder. Carefully add the water to the beaker. **Record** the volume of water used on the data table 1.
4. Allow the system of water and crystals to stand for 2 minutes. **Record** your observations in the data table 2.
5. Then gently stir the crystals and water. This change results in a solution being formed. **Record** the observations of the solution on data table 2.
6. Using a thermometer find the temperature of the copper II chloride solution. **Record** the temperature on the data table 1.
7. Obtain a piece of aluminum (Al) foil from the teacher.  **Find** the mass of the foil. **Record** the mass of the aluminum foil on the data table 1.
8. Place the loosely rolled aluminum foil into the solution. **Record** what happens between the aluminum foil and copper II chloride solution on data table 2.
9. Using the thermometer as a stirring stick, gently stir the mixture. Measure the temperature of the mixture when it reaches the highest value. **Record** the final temperature on the data table 1.
10. After the investigation is completed, depose of the end products as directed by the teacher. Return all the clean equipment and materials back to the appropriate places.

Student

Handout

**Observing Evidence of Chemical Interaction**

**Lab Write Up**

**Purpose:** Write out a purpose for this investigation.

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**Data Table:**

Table 1 – Quantitative Observations

|  |  |
| --- | --- |
| Mass of copper II chloride  |  |
| Volume of distilled water |  |
| Mass of aluminum foil |  |
| Temperature before mixing |  |
| Temperature after mixing |  |
| Change in temperature |  |

Table 2 – Qualitative Observations

|  |
| --- |
| The color of copper II chloride is |
| The color of water and copper II chloride is |
| The color of copper II chloride solution is |
| The color of mixture of Al foil and CuCl2 is |

**Discussion:** Answer the questions in a complete sentence.

1. When water and copper II chloride were mixed, how many phases did the resulting solution have?

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1. Was the mixture of water and copper II chloride a chemical or physical change? Why?

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1. Does the temperature increase or decrease as result of mixing the Al foil with the copper II chloride solution?

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1. Was the mixture of aluminum foil and CuCl2 solution a chemical or physical change? Why?

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1. What evidence could you use to indicate there was an interaction between the Al and CuCl2 solution in question 4?

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1. Al + CuCl2 🡪 AlCl3 + Cu is the chemical equation for this reaction. What are the reactants? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ What are the products? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. Balance the chemical equation for this reaction.

\_\_\_\_Al + \_\_\_\_CuCl2 🡪 \_\_\_\_\_AlCl3 + \_\_\_\_\_Cu

**Conclusion**: In at least 3 complete sentences, explain what you learned or did in this lab.

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