**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_**

**CLASSIFYING TYPES OF CHEMICAL REACTIONS**

**FLOW CHART**

DOUBLE REPLACEMENT:

An element in two compounds switch places

AB + CD 🡪 AD +CB

Ex. MgO + BeS 🡪 MgS + BeO

SINGLE REPLACEMENT:

An element takes the place of an element in a compound

A + BC 🡪 B +AC

Ex. Mg + BeO 🡪 Be + MgO

SYNTHESIS:

Reactants combine to make a single compound

A + B 🡪 AB

Ex. Mg + O 🡪 MgO

COMBUSTION:

A Hydrocarbon and Oxygen give you water and carbon dioxide

CH + O2 🡪H2O + CO2

Ex. C6H12 + O2 🡪 H2O + CO2

DECOMPOSITION:

A reactant breaks up into pieces

AB 🡪 A + B

Ex. MgO 🡪 Mg + O

Do you have 2 compounds as the reactants?

Do you have 1 element and 1 compound as the reactants?

Are H2O and CO2 the only 2 products?

One Product?

One Reactant?

TYPES OF REACTIONS

QUESTIONS TO ASK: